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~~solution is a~~

~~homogeneous~~

~~mixture of two or~~

~~9. more chemically~~

~~non-reacting~~

~~substances. The~~

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Components of a solution generally cannot be separated by filtration, settling or centrifuging. 2. A solution may be classified as solid, liquid or a gaseous solution. 3.

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1. Solution

Chemistry Notes. 2.

Solutions • Basic
definitions -

Solution:

homogenous

mixture - Solute:

substance that is

dissolved - Solvent:

substance that

contains the solute

in a solution •

Process of

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Dissolving: –
Molecules of the
solvent “pull apart”
the solute

molecules, and
distributes them
within the solvent •

How to tell the
solute from the
solvent – The
substance that is in
the greater amount
is the solvent – The
substance that

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keeps its phase is
the solvent.

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Chapter 2 Solutions
Solution is a
homogeneous
mixture of two or
more substances in
same or different
physical phases.

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The substances forming the solution are called components of the solution. On the basis of number of components a solution of two components is called binary solution. Solute and Solvent

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Solutions: Solutions

are the

homogeneous

mixtures of two or

more than two

components.

Binary solution: A

solution having two

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components is called a binary solution.

Components of a binary solution. It includes solute and solvent. When ...

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Solutions 1. In manufacture of soft drinks and soda water, CO_2 is passed at high pressure to increase its solubility. 2. To minimise the painful effects (bends) accompanying the decompression of deep sea divers. O

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2 diluted with
less... 3. At high
altitudes, the ...

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Notes.
Classification.
Solutions which
contain two

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components in it are called Binary Solutions.

Substances which are used to prepare a solution are called as Components. The component that is present in the largest quantity is known as Solvent. Solvent determines the physical state

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in which solution
exists.

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Types of solutions,
expression of
concentration of
solutions of solids
in liquids, solubility
of gases in liquids,
solid solutions,

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Colligative
properties - relative
lowering of vapour
pressure, Raoult's
law, elevation of
boiling point,
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conductance in
electrolytic
solutions, specific
and molar
conductivity
variation of

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conductivity with concentration, Kohlrausch's Law, electrolysis and Laws of electrolysis (elementary idea), dry cell- electrolytic cells and Galvanic cells; lead accumulator, EMF of a cell, standard electrode potential, Relation between Gibbs energy

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Change and EMF of
a cell, fuel cells;
corrosion.

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Solution: A solution
is a homogeneous
mixture of two (or
more) substances,
the composition of

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which may vary between certain limits. A solution consisting of two components is called binary solution. The component which is present in large quantity is called solvent and the component which is small in quantity is called solute.

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