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Chemistry comes alive with Chemistry Matters for GCE ' O ' Level (2nd edition). While fulfilling the latest GCE ' O ' Level Chemistry syllabus of the Ministry of Education, Singapore, this textbook also engages, inspires thoughtful learning and cultivates essential 21st Century Competencies. • Makes Chemistry engaging, clear and simple

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Chemistry Matters for GCE 'O' Level (2nd Edition): Full Solutions to Textbook Questions Chapter 16 © 2013 Marshall Cavendish International (Singapore) Private Limited 16.3 (c) The nuclear disaster has caused radioactive substances to spread into the atmosphere and pollute the air.

Chapter 16 The Periodic Table - O Level Chemistry tuition ...

The error in print is on Page 271 under the Chapter of “ Electrolysis ” in Chemistry Matters for Pure Chemistry syllabus. In the Key Ideas which acts like sub-section summary, it states that “ OH⁻ ions will be discharged as oxygen gas. However, if a solution is concentrated in halide ions, the halogen gas will be produced. ”

Error in Marshall Cavendish GCE O-Level Chemistry Matters ...

Chemistry Matters for GCE 'O' Level (2nd Edition): Textbook Answers Chapter 9 © 2013 Marshall Cavendish International (Singapore) Private Limited 9.2 (c) Number of moles of C₁₂ gas = mass of C₁₂ / molar mass of C₁₂ = 1.4271 / 71 = 0.02 mol (d) Number of moles of sodium chloride = number of ions of sodium chloride / 6 × 10²³ = 6 × 10²² / 6 × 10²³ = 0.1 mol

CM_TB answers_C09.pdf - nd Chemistry Matters for GCE ...

Test Yourself 4.1 (page 74) 1. (Any four differences) Mass Weight • An amount of matter • A gravitational force • A scalar quantity (i.e. has only magnitude) • A vector quantity (i.e. has both magnitude and direction) • SI unit: kilogram (kg) • SI

(PDF) Physics Matters for GCE 'O' Level (4th Edition ...

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Chemistry Matters for GCE ' O ' Level (2 ndEdition): Textbook Answers Chapter 10 © 2013 Marshall Cavendish International (Singapore) Private Limited 10.2 Test Yourself 10.2 (page 167) 1. $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$ Volume of carbon dioxide formed = 2 ×

[CM_TB answers_C10.pdf - nd Chemistry Matters for GCE ...](#)

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