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39 E On this test, the easiest way to balance an equation is by plugging in the answers. The test writer tells you that the coefficient in front of H_2O is 6, so put that in there. Now, see which of the answer choices, if placed in front of NF_3 , would result in a balanced equation. If, for instance, we try choice (B), we'd have 1 mole of N on the left, which would give us 3 moles of F on the left.

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Answers 1. $1 \text{ SnO}_2 + 2 \text{ H}_2 \rightarrow 1 \text{ Sn} + 2 \text{ H}_2\text{O}$ 2. $3 \text{ KOH} + 1 \text{ H}_3\text{PO}_4 \rightarrow 1 \text{ K}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ 3. $2 \text{ KNO}_3 + 1 \text{ H}_2\text{CO}_3 \rightarrow 1 \text{ K}_2\text{CO}_3 + 2 \text{ HNO}_3$ 4. $1 \text{ Na}_3\text{PO}_4 + 3 \text{ HCl} \rightarrow 3 \text{ NaCl} + 1 \text{ H}_3\text{PO}_4$ 5. $1 \text{ TiCl}_4 + 2 \text{ H}_2\text{O} \rightarrow 1 \text{ TiO}_2 + 4 \text{ HCl}$ 6. $1 \text{ C}_2\text{H}_6\text{O} + 3 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 3 \text{ H}_2\text{O}$ 7. $2 \text{ Fe} + 6 \text{ HC}_2\text{H}_3\text{O}_2 \rightarrow 2 \text{ Fe}(\text{C}_2\text{H}_3\text{O}_2)_3 + 3 \text{ H}_2$ 8. $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$ 9. $1 \text{ B}_2\text{Br}_6 + 6 \text{ HNO}_3 \rightarrow 2 \text{ B}(\text{NO}_3)_3 + 6 \text{ HBr}$ 10. $4 \text{ NH}_4\text{OH} + 1 \text{ K}_2\text{Al}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O} \rightarrow 1 \text{ Al}(\text{OH})_3 + 2 (\text{NH}_4)_2\text{SO}_4 + \dots$

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Quiz: Chemistry For Grade 10 Student
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