

Chapter 4 Review Arrangement Of Electrons In Atoms

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CHAPTER 4 REVIEW Arrangement of Electrons in Atoms SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two electrons in an atom may have the

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~~Chapter 4 Arrangement Of Electrons In Atoms Mixed Review~~

CHAPTER 4 REVIEW . Arrangement of Electrons in Atoms . SHORT ANSWER Answer the following questions in the space provided. 1. d-- How many quantum numbers are used to describe the properties of electrons in atomic orbitals? (a) 1 (c) 3 (b) 2 (d) 4 . a,-- A spherical electron cloud surrounding an atomic nucleus would best represent

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CHAPTER 4 REVIEW Arrangement of Electrons in Atoms states that a maximum of two electrons can occupy a single atomic orbital but only if the electrons have opposite spins Hund's rule states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins can occupy the same orbitals

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